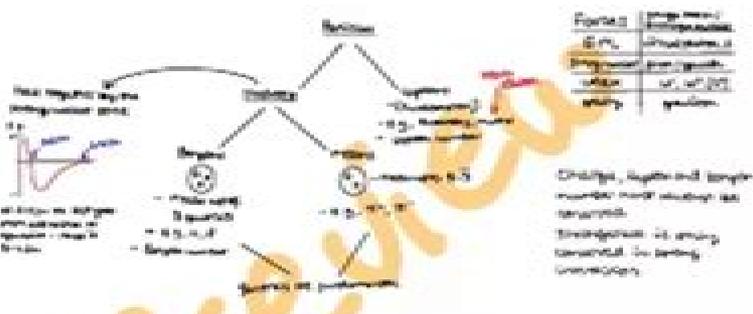


Particles and Quantum



To convert mass to energy, use $E=mc^2$
 Annihilation - pair of particles and its corresponding antiparticle collide and their rest energy is converted into photons



Pair production - when a photon turns into a particle and its corresponding antiparticle. The photon must have at least the same amount of energy as the rest energy of the particles



Excitation

Exciting means they gain energy through energy to climb their atomic energy levels - $1st$ level
 Alpha particles - Helium nucleus, $2He$
 - Highly ionising, so it is dangerous
 - Relatively penetrating - stopped by a plate of paper or foil in air
 ${}_{2}^{4}He \rightarrow {}_{2}^{3}He + {}_{0}^{1}H$
 Beta - fast moving electron $\rightarrow e^-$
 - Medium penetrating and relatively quickly - stopped by a few cm of aluminium
 - Beta decay - ${}_{Z}^{A}X \rightarrow {}_{Z-1}^{A}Y + e^- + \bar{\nu}_e$ antineutrino to conserve spin/momentum



Gamma - high energy EM ray or photon emitted from a high energy nucleus $\rightarrow \gamma$
 - Relatively penetrating
 - Highly penetrating - stopped by a few m of concrete

Isotopes - same element - same number of protons
 - different relative atomic mass - different number of neutrons
 ${}_{Z}^{A}X \rightarrow {}_{Z}^{A-1}X + e^- + \bar{\nu}_e$

Nuclear charge - $\frac{Q}{m} [C/kg]$
 - charge to mass ratio

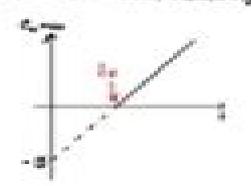
Electron volt - charge of an electron when accelerated through a pot of one volt
 - $1.6 \times 10^{-19} J$
 - eV



Photoelectric effect

Use photons of sufficient energy on electrons on the surface of the metal, liberating them

- $E_{photon} = hf = \phi + E_{max}$
- 1 - photon absorbed - hf (total hf) energy of the photon
- 2 - frequency
- 3 - work function - minimum energy to liberate electrons
- 4 - E_{max} - energy that left after ϕ is subtracted
- 5 - stopped frequency - minimum frequency needed for electrons to be liberated



The photoelectric effect proved the particle nature of light due to the fact that the maximum kinetic energy of electrons E_k is not linearly dependent on the wavelength but is linearly dependent on the frequency of light. Below a certain limit f_0 no electrons are emitted but as wavelength of light is increased the number of electrons that are emitted per second

Stopping potential - the pot. needed to stop any electrons from reaching the other side

- $E_{photon} = hf = \phi + E_{max}$
- $E_{max} = eV_s$ - charge of electron
- V_s - stopping potential

De Broglie

The de Broglie wavelength is the wavelength that a particle can have. The wave nature of a particle was proven by diffraction of a crystal lattice which produces similar groups that propagate across space. This is because the electron is diffracting within the crystal lattice



$\lambda = \frac{h}{p} = \frac{h}{mv}$
 $\lambda = \frac{h}{m\lambda v}$
 $f = \text{frequency / wave velocity}$

Fluorescent tubes



This tube contains mercury under a high voltage is applied. This excites the gas molecules to which some of the mercury atoms emit photons when they relax. When the gas relaxes it emits UV radiation. This UV radiation is higher energy than what the electrons have, so they produce in the tube light. A phosphor coating in the tube absorbs the photons and emit visible light

Energy levels

Electrons can be excited to higher energy levels either by absorbing photons from the light or by the collision. In a discharge tube or in a gas discharge cell when ionizing the electron energy

Ionization energy - is enough energy to pull an electron out of an atom, so will completely change the atom



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N Colangelo



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Table of Contents A Level Physics Notes

1. Understanding the eBook A Level Physics Notes
 - The Rise of Digital Reading A Level Physics Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying A Level Physics Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an A Level Physics Notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from A Level Physics Notes
 - Personalized Recommendations
 - A Level Physics Notes User Reviews and Ratings
 - A Level Physics Notes and Bestseller Lists

5. Accessing A Level Physics Notes Free and Paid eBooks
 - A Level Physics Notes Public Domain eBooks
 - A Level Physics Notes eBook Subscription Services
 - A Level Physics Notes Budget-Friendly Options
6. Navigating A Level Physics Notes eBook Formats
 - ePub, PDF, MOBI, and More
 - A Level Physics Notes Compatibility with Devices
 - A Level Physics Notes Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of A Level Physics Notes
 - Highlighting and Note-Taking A Level Physics Notes
 - Interactive Elements A Level Physics Notes
8. Staying Engaged with A Level Physics Notes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers A Level Physics Notes
9. Balancing eBooks and Physical Books A Level Physics Notes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection A Level Physics Notes
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine A Level Physics Notes
 - Setting Reading Goals A Level Physics Notes
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of A Level Physics Notes
 - Fact-Checking eBook Content of A Level Physics Notes
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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