

## VOLUMETRIC ANALYSIS -CLASS XI

### EXPERIMENT-1

#### Aim:

To prepare 100ml  $\frac{M}{20}$  sodium carbonate solution and determine the Molarity and the strength of the given solution of hydrochloric acid using  $\frac{M}{20}$  sodium carbonate solution.

#### Apparatus required:

Burette, Pipette, Burette stand, conical flask.

#### Chemicals Required:

Sodium carbonate, hydrochloric acid, Indicator (methyl orange), distilled water.

#### Theory:

- To prepare 100ml  $\frac{M}{20}$  sodium carbonate solution

Weight of  $\text{Na}_2\text{CO}_3$  required for preparing 100 ml  $M/20$   $\text{Na}_2\text{CO}_3$  solution.

$$\frac{M}{20} = \frac{\text{weight of sodium carbonate required for 100 ml} \times 1000}{\text{molar mass of sodium carbonate} \times 100}$$

$$\therefore \frac{1}{20} = \frac{x \times 1000}{106 \times 100} \Rightarrow \frac{x \times 10}{106}$$

$$\frac{1}{20} = \frac{10x}{106}$$

$$= \frac{106}{200} = 0.53 \text{ gm.}$$

# Xi Chemistry Practical Procedure Volumetric Analysis

**C. Remigius Fresenius**



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