

## Calculating Average Atomic Mass Worksheet:

- 1) Three isotopes of Silicon occur in nature:

| Isotopes of Silicon: | Percent Abundance: | Atomic Mass: |
|----------------------|--------------------|--------------|
| Silicon-28           | 92.23%             | 27.97693 amu |
| Silicon-29           | 4.68%              | 28.97649 amu |
| Silicon-30           | 3.09%              | 29.97377 amu |

Calculate the average atomic mass for the three isotopes of Silicon.

$$(27.97693 \times 0.9223) + (28.97649 \times 0.0468) + (29.97377 \times 0.0309) = 28.0854 \text{ am}$$

Si-28                      Si-29                      Si-30

- 2) Two isotopes of Rubidium occur naturally:

| Isotopes of Rubidium: | Percent Abundance: | Atomic Mass: |
|-----------------------|--------------------|--------------|
| Rubidium-85           | 72.15%             | 84.9118 amu  |
| Rubidium-87           | 27.85%             | 86.9092 amu  |

Calculate the average atomic mass for the two isotopes of Rubidium.

$$(84.9118 \times 0.7215) + (86.9092 \times 0.2785) = 85.468$$

Rubidium-85                      Rubidium-87

- 3) Only two isotopes of Copper naturally occur:

| Isotopes of Copper: | Percent Abundance: | Atomic Mass: |
|---------------------|--------------------|--------------|
| Copper-63           | 69.17%             | 62.9396 amu  |
| Copper-65           | 30.83%             | 64.9278 amu  |

Calculate the average atomic mass for the two isotopes of Copper.

$$(62.9396 \times 0.6917) + (64.9278 \times 0.3083) = 63.552$$

Copper-63                      Copper-65

# Calculating Average Atomic Mass Worksheet Answers

**David Baud**



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