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|----|---|---|
| 1. | Which of the following best describes what causes water molecules to exhibit properties such as cohesive and adhesive behavior? | Water molecules are polar, having partially positive and partially negative ends. |
| 2. | Which statement correctly compares the pH of water to that of other substances? | The pH of water is neutral; it is higher than the pH of acids and lower than the pH of bases. |
| 3. | The slight negative charge at one end of one water molecule is attracted to the slight positive charge of another water molecule. What is this attraction called? | A hydrogen bond |
| 4. | Which of the following best explains how water is able to dissolve ionic substances? | The positive and negative charges of the water molecule pull on the negative and positive ends of the ionic molecule to break them apart. |
| 5. | During winter's freezing temperatures, some plants can remove water from their cells to reduce the amount of water that will freeze inside them. Why would plants need to do this? | The expansion of excess water as it turns to ice inside the cells would break the cells apart. |
| 6. | Which of the following best describes how the density of water differs from that of other substances? | The density of water depends on its temperature, unlike other substances. |

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