

### Basic Stoichiometry Post-Lab Homework Exercises

1. Load the "Reactants, Products, and Leftovers" simulation and work through each of the levels of the Game! At <http://phet.colorado.edu/> or googling "phet." You may have to download or update Java on your computer.

**Show your work for all calculations. Remember: Labels, units, sig figs and cancellation!**

2. For the reaction  $P_4 + 6Cl_2 \rightarrow 4PCl_3$ , determine how many moles of chlorine  $Cl_2$  would be needed to react with 4.0 moles of phosphorus  $P_4$  to entirely use up all the phosphorus. 4) \_\_\_\_\_

3. If 3.00 moles of  $P_4$  reacted with 22.0 moles  $Cl_2$  according to the above reaction, determine:

- a) How many moles  $PCl_3$  are produced a) \_\_\_\_\_
- b) How many moles of  $P_4$  are left in excess after the reaction (if any) b) \_\_\_\_\_
- c) How many moles of  $Cl_2$  are left in excess after the reaction (if any) c) \_\_\_\_\_

In reality, reactants don't have to react in perfect whole-numbers of moles. In a two-reactant synthesis reaction, usually **one reactant gets entirely used up (and determines how much product is made)**, even if that means using fractions of a mole of reactant. For instance, when solid, metallic aluminum Al and red, liquid bromine  $Br_2$  are brought together, they make a white solid according to the reaction  $2Al + 3Br_2 \rightarrow 2AlBr_3$ . If 5.0 moles of aluminum Al was reacted with 10 moles bromine  $Br_2$ , all five moles of aluminum would react, with only 7.5 moles bromine. (2:3 mole ratio) This would produce only 5.0 moles of  $AlBr_3$ , leaving 2.5 moles of excess  $Br_2$  behind.

5. Now assume 3.0 moles Al and 4.0 moles  $Br_2$  react.

- a) How much (in moles)  $AlBr_3$  gets produced? a) \_\_\_\_\_
- b) Which chemical must be the *excess reactant*? b) \_\_\_\_\_
- c) Which chemical is the *limiting reactant*? c) \_\_\_\_\_
- d) If all the limiting reactant gets used up, how much of the excess reactant is left? d) \_\_\_\_\_

# Basic Stoichiometry Post Lab Homework Exercises Answers

**JA Banks**



## **Basic Stoichiometry Post Lab Homework Exercises Answers:**

*Lab Manual for General, Organic, and Biochemistry* Denise Guinn, Rebecca Brewer, 2009-08-21 Teaching all of the necessary concepts within the constraints of a one term chemistry course can be challenging Authors Denise Guinn and Rebecca Brewer have drawn on their 14 years of experience with the one term course to write a textbook that incorporates biochemistry and organic chemistry throughout each chapter emphasizes cases related to allied health and provides students with the practical quantitative skills they will need in their professional lives *Essentials of General Organic and Biochemistry* captures student interest from day one with a focus on attention getting applications relevant to health care professionals and as much pertinent chemistry as is reasonably possible in a one term course Students value their experience with chemistry getting a true sense of just how relevant it is to their chosen profession To browse a sample chapter view sample ChemCasts and more visit [www.whfreeman.com/gob](http://www.whfreeman.com/gob) *Working with Chemistry* Donald J. Wink, Sharon Fetzer-Gislason, Julie Ellefson Kuehn, 2004-02-20 With this modular laboratory program students build skills using important chemical concepts and techniques to the point where they are able to design a solution to a scenario drawn from a professional environment The scenarios are drawn from the lives of people who work with chemistry every day ranging from field ecologists to chemical engineers and include many health professionals as well *Carolina Science and Math* Carolina Biological Supply Company, 2003 [Bibliography of Agriculture with Subject Index](#) ,1993-10

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## **Table of Contents Basic Stoichiometry Post Lab Homework Exercises Answers**

1. Understanding the eBook Basic Stoichiometry Post Lab Homework Exercises Answers
  - The Rise of Digital Reading Basic Stoichiometry Post Lab Homework Exercises Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Basic Stoichiometry Post Lab Homework Exercises Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Basic Stoichiometry Post Lab Homework Exercises Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Basic Stoichiometry Post Lab Homework Exercises Answers

- Personalized Recommendations
  - Basic Stoichiometry Post Lab Homework Exercises Answers User Reviews and Ratings
  - Basic Stoichiometry Post Lab Homework Exercises Answers and Bestseller Lists
5. Accessing Basic Stoichiometry Post Lab Homework Exercises Answers Free and Paid eBooks
- Basic Stoichiometry Post Lab Homework Exercises Answers Public Domain eBooks
  - Basic Stoichiometry Post Lab Homework Exercises Answers eBook Subscription Services
  - Basic Stoichiometry Post Lab Homework Exercises Answers Budget-Friendly Options
6. Navigating Basic Stoichiometry Post Lab Homework Exercises Answers eBook Formats
- ePub, PDF, MOBI, and More
  - Basic Stoichiometry Post Lab Homework Exercises Answers Compatibility with Devices
  - Basic Stoichiometry Post Lab Homework Exercises Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Basic Stoichiometry Post Lab Homework Exercises Answers
  - Highlighting and Note-Taking Basic Stoichiometry Post Lab Homework Exercises Answers
  - Interactive Elements Basic Stoichiometry Post Lab Homework Exercises Answers
8. Staying Engaged with Basic Stoichiometry Post Lab Homework Exercises Answers
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Basic Stoichiometry Post Lab Homework Exercises Answers
9. Balancing eBooks and Physical Books Basic Stoichiometry Post Lab Homework Exercises Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Basic Stoichiometry Post Lab Homework Exercises Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Basic Stoichiometry Post Lab Homework Exercises Answers
- Setting Reading Goals Basic Stoichiometry Post Lab Homework Exercises Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Basic Stoichiometry Post Lab Homework Exercises Answers

- Fact-Checking eBook Content of Basic Stoichiometry Post Lab Homework Exercises Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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