

## Balancing Chemical Equations Worksheet

1.  $\underline{2} \text{ H}_2 + \underline{\quad} \text{ O}_2 \rightarrow \underline{2} \text{ H}_2\text{O}$
2.  $\underline{\quad} \text{ N}_2 + \underline{3} \text{ H}_2 \rightarrow \underline{2} \text{ NH}_3$
3.  $\underline{\quad} \text{ S}_8 + \underline{12} \text{ O}_2 \rightarrow \underline{8} \text{ SO}_3$
4.  $\underline{2} \text{ N}_2 + \underline{\quad} \text{ O}_2 \rightarrow \underline{2} \text{ N}_2\text{O}$
5.  $\underline{2} \text{ HgO} \rightarrow \underline{2} \text{ Hg} + \underline{\quad} \text{ O}_2$
6.  $\underline{4} \text{ CO}_2 + \underline{6} \text{ H}_2\text{O} \rightarrow \underline{\quad} \text{ C}_6\text{H}_{12}\text{O}_6 + \underline{6} \text{ O}_2$
7.  $\underline{\quad} \text{ Zn} + \underline{2} \text{ HCl} \rightarrow \underline{\quad} \text{ ZnCl}_2 + \underline{\quad} \text{ H}_2$
8.  $\underline{\quad} \text{ SiCl}_4 + \underline{4} \text{ H}_2\text{O} \rightarrow \underline{\quad} \text{ H}_4\text{SiO}_4 + \underline{4} \text{ HCl}$
9.  $\underline{2} \text{ Na} + \underline{2} \text{ H}_2\text{O} \rightarrow \underline{2} \text{ NaOH} + \underline{\quad} \text{ H}_2$
10.  $\underline{2} \text{ H}_3\text{PO}_4 \rightarrow \underline{\quad} \text{ H}_4\text{P}_2\text{O}_7 + \underline{\quad} \text{ H}_2\text{O}$
11.  $\underline{\quad} \text{ C}_{10}\text{H}_{16} + \underline{8} \text{ Cl}_2 \rightarrow \underline{10} \text{ C} + \underline{16} \text{ HCl}$
12.  $\underline{\quad} \text{ CO}_2 + \underline{2} \text{ NH}_3 \rightarrow \underline{\quad} \text{ OC}(\text{NH}_2)_2 + \underline{\quad} \text{ H}_2\text{O}$
13.  $\underline{4} \text{ Si}_2\text{H}_3 + \underline{17} \text{ O}_2 \rightarrow \underline{8} \text{ SiO}_2 + \underline{6} \text{ H}_2\text{O}_3$
14.  $\underline{2} \text{ Al}(\text{OH})_3 + \underline{3} \text{ H}_2\text{SO}_4 \rightarrow \underline{\quad} \text{ Al}_2(\text{SO}_4)_3 + \underline{6} \text{ H}_2\text{O}$
15.  $\underline{4} \text{ Fe} + \underline{3} \text{ O}_2 \rightarrow \underline{2} \text{ Fe}_2\text{O}_3$
16.  $\underline{\quad} \text{ Fe}_2(\text{SO}_4)_3 + \underline{6} \text{ KOH} \rightarrow \underline{3} \text{ K}_2\text{SO}_4 + \underline{2} \text{ Fe}(\text{OH})_3$
17.  $\underline{2} \text{ C}_7\text{H}_6\text{O}_2 + \underline{15} \text{ O}_2 \rightarrow \underline{14} \text{ CO}_2 + \underline{6} \text{ H}_2\text{O}$
18.  $\underline{\quad} \text{ H}_2\text{SO}_4 + \underline{8} \text{ HI} \rightarrow \underline{\quad} \text{ H}_2\text{S} + \underline{4} \text{ I}_2 + \underline{4} \text{ H}_2\text{O}$
19.  $\underline{4} \text{ FeS}_2 + \underline{11} \text{ O}_2 \rightarrow \underline{2} \text{ Fe}_2\text{O}_3 + \underline{8} \text{ SO}_2$
20.  $\underline{2} \text{ Al} + \underline{3} \text{ FeO} \rightarrow \underline{\quad} \text{ Al}_2\text{O}_3 + \underline{3} \text{ Fe}$
21.  $\underline{\quad} \text{ Fe}_2\text{O}_3 + \underline{3} \text{ H}_2 \rightarrow \underline{2} \text{ Fe} + \underline{3} \text{ H}_2\text{O}$
22.  $\underline{\quad} \text{ Na}_2\text{CO}_3 + \underline{2} \text{ HCl} \rightarrow \underline{2} \text{ NaCl} + \underline{\quad} \text{ H}_2\text{O} + \underline{\quad} \text{ CO}_2$
23.  $\underline{2} \text{ K} + \underline{\quad} \text{ Br}_2 \rightarrow \underline{2} \text{ KBr}$
24.  $\underline{\quad} \text{ C}_7\text{H}_{16} + \underline{11} \text{ O}_2 \rightarrow \underline{7} \text{ CO}_2 + \underline{8} \text{ H}_2\text{O}$
25.  $\underline{2} \text{ P}_4 + \underline{5} \text{ O}_2 \rightarrow \underline{2} \text{ P}_2\text{O}_5$

# Balancing Chemical Equations Answer Sheet

**Camille Linda Wainwright**



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