

STOICHIOMETRY PROBLEMS

1. The combustion of a sample of butane, C_4H_{10} (lighter fluid), produced 2.46 grams of water.



- (a) How many moles of water formed?
- (b) How many moles of butane burned?
- (c) How many grams of butane burned?
- (d) How much oxygen was used up in moles?
- (e) How much oxygen was used up in grams?

2. Sodium metal burns in air according to the balanced reaction below.



- (a) How many moles of oxygen are needed to react with 9.5 g of sodium completely?
- (b) How many grams of sodium are needed to produce 12.5 g of sodium oxide?

3. In the combustion of 54.5 g of butane (C_4H_{10}), how many grams of CO_2 are produced? Write and balance the equation before solving.

Stoichiometry 8b Extra Practice Problems Answers

Christian Drosten



Stoichiometry 8b Extra Practice Problems Answers:

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In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Stoichiometry 8b Extra Practice Problems Answers**," an enthralling opus penned by a highly acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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web the iodination of acetone is convenient because the color of the iodine reactant is readily observable and the change in concentration of I_2 the iodine can be easily monitored by our spectrophotometer and the reaction proceeds quickly enough for multiple trials but slowly enough to obtain measurements

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web i the iodination of acetone the rate of this reaction is found to depend on the concentration of hydrogen ion in the solution as well as presumably on the concentrations of the two reactants

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web video answer the data was given by an point we arrange the dates in ascending order this is the pointing we need to build a stem and plot graph to present quantity data included a graphical input this is the stem part and this is the leaf item yours is

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