

In Table 4 below, $[\text{Fe}^{3+}]$ after mixing is obtained by using $M_1V_1 = M_2V_2$. This data is to be transferred to Table 5 below as the initial $[\text{Fe}^{3+}]$ (column 2).

Similarly in Table 4 of Data sheet $[\text{SCN}^-]$ after mixing obtained using $M_1V_1 = M_2V_2$. This data is then to be transferred to Table 5 below as the initial $[\text{SCN}^-]$ (column 3).

Table 4. Initial Concentration Data

Test Tube	$[\text{Fe}^{3+}]$ Before Mixing	$[\text{Fe}^{3+}]$ After Mixing ¹ (Initial $[\text{Fe}^{3+}]$)	$[\text{SCN}^-]$ Before Mixing	$[\text{SCN}^-]$ After Mixing ¹ (Initial $[\text{SCN}^-]$)
F	0.01	0.0025M	0.00113	2.825×10^{-4}
G	0.01	0.0025M	0.00113	5.65×10^{-4}
H	0.01	0.005M	0.00113	2.825×10^{-4}
I	0.01	0.005M	0.00113	1.13×10^{-4}

(Show calculations for initial concentration data on separate sheet)

In Table 5 below, Equilibrium $[\text{FeSCN}^{2+}]$ is obtained from the Calibration line using absorbance values of solutions F through I.

Equilibrium $[\text{Fe}^{3+}] = \text{Initial } [\text{Fe}^{3+}] - \text{Equilibrium } [\text{FeSCN}^{2+}]$

Equilibrium $[\text{SCN}^-] = \text{Initial } [\text{SCN}^-] - \text{Equilibrium } [\text{FeSCN}^{2+}]$

Table 5. Equilibrium Constant Calculation

Test Tube	Initial $[\text{Fe}^{3+}]$	Initial $[\text{SCN}^-]$	Equilibrium $[\text{FeSCN}^{2+}]$ (from graph)	Equilibrium $[\text{Fe}^{3+}]$	Equilibrium $[\text{SCN}^-]$	K_c
F	0.0025	2.825×10^{-4}				
G	0.0025	5.65×10^{-4}				
H	0.005	2.825×10^{-4}				
I	0.005	1.13×10^{-4}				

(Show calculations for Equilibrium Constant on separate sheet)

Average Equilibrium Constant _____

Determining Equilibrium Constant Lab Report Answers

Katrin Zwirgmaier



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