

Date Performed: 10/26/09

Date Submitted: 11/02/09

Instructor:

Name:

Partner:

Acid-Base Titrations

Objective

- To standardize a sodium hydroxide solution (NaOH) with a hydrochloric acid solution (HCl) of known concentration by the reaction.
- To determine the molar mass of an unknown, monoprotic acid
- To determine the percentage of Acetylsalicylic Acid in a tablet of aspirin.

Experimental Data

Part A-Standardization of Sodium Hydroxide Solution

Titration	Vol NaOH (mL)
1	20.00
2	20.21

Part B-Molar Mass of an Unknown Monoprotic Acid

Sample	Mass Beaker (g)	Mass Sample + Beaker (g)	Vol NaOH (mL)
1	119.50	120.50	20.95
2	88.32	89.32	19.95

Part C-Percentage of Acetylsalicylic Acid in a Commercial Aspirin Tablet

Mass Beaker (g)	Mass Tablet +Beaker (g)	Vol NaOH (mL)
122.08	122.46	8.49

Sample Calculations

1) Molarity NaOH

(Molarity HCl)(Volume HCl)=(Molarity NaOH)(Volume NaOH)

(0.200M)(25.00mL)=(Molarity NaOH)(20.00mL)

5.00=X(20.00)

Trail 1: Molarity of NaOH=0.250M NaOH

Trail 2: Molarity of NaOH=0.247M NaOH

Average=0.249M NaOH

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Jeffrey Stanton



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Integrated Approach to Coordination Chemistry Rosemary A. Marusak, Kate Doan, Scott D. Cummings, 2007-03-07
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Hulanicki,1987 **Aqueous Acid-base Equilibria and Titrations** Robert De Levie,1999 This book will give students a thorough grounding in pH and associated equilibria material absolutely fundamental to the understanding of many aspects of chemistry It is in addition a fresh and modern approach to a topic all too often taught in an outmoded way This book uses new theoretical developments which have led to more generalized approaches to equilibrium problems these approaches are often simpler than the approximations which they replace Acid base problems are readily addressed in terms of the proton condition a convenient amalgam of the mass and charge constraints of the chemical system considered The graphical approach of Bjerrum Hagg and Sillen is used to illustrate the orders of magnitude of the concentrations of the various species involved in chemical equilibria Based on these concentrations the proton condition can usually be simplified often leading directly to the value of the pH In the description of acid base titrations a general master equation is developed It provides a continuous and complete description of the entire titration curve which can then be used for computer based comparison with experimental data Graphical estimates of the steepness of titration curves are also developed from which the practicality of a given titration can be anticipated Activity effects are described in detail including their effect on titration curves The discussion emphasizes the distinction between equilibrium constants and electrometric pH measurements which are subject to activity corrections and balance equations and spectroscopic pH measurements which are not Finally an entire chapter is devoted to what the pH meter measures and to the experimental and theoretical uncertainties involved

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