

Prentice Hall Physical Science

Chapter 4 Atomic Structure

4.1 Studying Atoms

- A. Ancient Greek Models of Atoms
 - Democritus believed all matter was made of extremely small particles that could not be divided
 - he called the particles atoms (from the Greek atomos meaning indivisible)
 - he thought there were different types of atoms with different properties (liquids had round, smooth atoms and solids were rough and prickly)
- B. Dalton's Atomic Theory
 - he proved atoms exist by discovering that compounds have a fixed composition
 - Dalton's Theory:
 - all elements are composed of atoms
 - all atoms of the same element have the same mass, and atoms of different elements have different masses
 - compounds contain atoms of more than one element
 - in a particular compound, atoms of different element always combine in the same way
 - he pictured atoms as solid spheres
 - in time it was found that not all of this was completely correct
- C. Thomson's Model of the Atom
 - some objects have a positive or negative charge and objects with the same charge repel and objects with opposite charges attract
 - Thomson used an electric current to learn more about atoms
 - he passed an electric current through a gas and a glowing beam appeared
 - the beam was repelled by a negatively charged metal plate (like charges repel and opposite charges attract)
 - Thomson concluded that the beam must be made of negatively charged particles, smaller than atoms.
 - since the atom was known to be electrically neutral, he developed a model in which negative charges were evenly scattered throughout an atom filled with a positively charged pudding like material
 - called the "Plum Pudding" model (or chocolate chip cookie dough)
- D. Rutherford's Atomic Theory
 - Rutherford discovered positively charged, fast moving particles called alpha particles
 - he shot some of these particles a piece of gold foil, thinking they would go straight through
 - when he actually performed the experiment, he found that the positively charged alpha particles were greatly deflected by the gold (figure 7, p. 104)
 - he determined that the alpha particles must have come very close to another positively charged particle causing them to be repelled and put off course
 - this led him to believe that the positive charge in an atom is not evenly distributed, but concentrated in a very small central area he called the nucleus
 - his model of the atom is a small, positively charged nucleus surrounded by negatively charged particles

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Table of Contents Chapter Assessment Prentice Hall Physical Science

1. Understanding the eBook Chapter Assessment Prentice Hall Physical Science
 - The Rise of Digital Reading Chapter Assessment Prentice Hall Physical Science
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Assessment Prentice Hall Physical Science
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter Assessment Prentice Hall Physical Science
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter Assessment Prentice Hall Physical Science
 - Personalized Recommendations
 - Chapter Assessment Prentice Hall Physical Science User Reviews and Ratings
 - Chapter Assessment Prentice Hall Physical Science and Bestseller Lists
5. Accessing Chapter Assessment Prentice Hall Physical Science Free and Paid eBooks
 - Chapter Assessment Prentice Hall Physical Science Public Domain eBooks
 - Chapter Assessment Prentice Hall Physical Science eBook Subscription Services

- Chapter Assessment Prentice Hall Physical Science Budget-Friendly Options
- 6. Navigating Chapter Assessment Prentice Hall Physical Science eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter Assessment Prentice Hall Physical Science Compatibility with Devices
 - Chapter Assessment Prentice Hall Physical Science Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter Assessment Prentice Hall Physical Science
 - Highlighting and Note-Taking Chapter Assessment Prentice Hall Physical Science
 - Interactive Elements Chapter Assessment Prentice Hall Physical Science
- 8. Staying Engaged with Chapter Assessment Prentice Hall Physical Science
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter Assessment Prentice Hall Physical Science
- 9. Balancing eBooks and Physical Books Chapter Assessment Prentice Hall Physical Science
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter Assessment Prentice Hall Physical Science
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter Assessment Prentice Hall Physical Science
 - Setting Reading Goals Chapter Assessment Prentice Hall Physical Science
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter Assessment Prentice Hall Physical Science
 - Fact-Checking eBook Content of Chapter Assessment Prentice Hall Physical Science
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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