

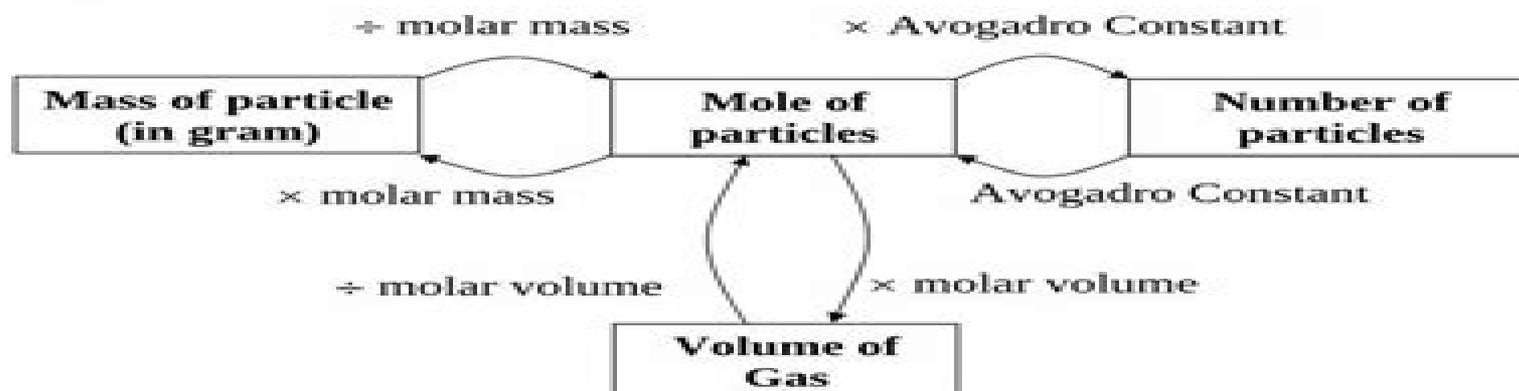
## Short Notes: Form 4 Chemistry

### Chemical Formulae and Equation

#### Calculation

<p><b>For Solid, liquid or gas</b></p> $\text{number of mole} = \frac{\text{mass of substance}}{\text{molar mass}}$ <p>Molar mass = RAM/RMM/RFM in gram</p>	<p><b>For gas (only)</b></p> $\text{number of mole} = \frac{\text{volume of gas}}{\text{molar volume}}$ <p>Molar volume = 24dm<sup>3</sup> at room temperature Molar volume = 22.4dm<sup>3</sup> at s.t.p.</p>
<p><b>For Solution</b></p> $\text{number of mole} = \frac{MV}{1000}$ <p>M = molarity V = Volume of solution in cm<sup>3</sup></p>	<p><b>For quantity of particle(atom,molecule,ion)</b></p> $\text{number of mole} = \frac{\text{quantity of particle}}{6.02 \times 10^{23}}$

#### Summary



# One School Short Notes Form 4 Chemistry

**R Barnett**



**One School Short Notes Form 4 Chemistry:**

## Reviewing **One School Short Notes Form 4 Chemistry**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**One School Short Notes Form 4 Chemistry**," an enthralling opus penned by a highly acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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