

11. Which statement about one atom of an element identifies the element?
- A) **The atom has 1 proton.**  
B) The atom has 2 neutrons.  
C) The sum of the number of protons and neutrons in the atom is 3.  
D) The difference between the number of neutrons and protons in the atom is 1.
12. Compared to a proton, an electron has
- A) a greater quantity of charge and the same sign  
B) a greater quantity of charge and the opposite sign  
C) the same quantity of charge and the same sign  
D) **the same quantity of charge and the opposite sign**
13. Every chlorine atom has
- A) 7 electrons  
B) 17 neutrons  
C) a mass number of 35  
D) **an atomic number of 17**
14. What can be determined if only the atomic number of an atom is known?
- A) the total number of neutrons in the atom, only  
B) the total number of protons in the atom, only  
C) the total number of protons and the total number of neutrons in the atom  
D) **the total number of protons and the total number of electrons in the atom**
15. Which of the following atoms has the greatest nuclear charge?
- A)  ${}_{7}^{14}\text{N}$    B)  ${}_{6}^{12}\text{C}$    C)  ${}_{7}^{11}\text{H}$    D)  ${}_{2}^{4}\text{He}$
16. What is the total number of neutrons in the nucleus of a neutral atom that has 19 electrons and a mass number of 39?
- A) 19   B) **20**   C) 39   D) 58
17. The weighted average of the atomic masses of the naturally occurring isotopes of an element is the
- A) **atomic mass of the element**  
B) atomic number of the element  
C) mass number of each isotope  
D) formula mass of each isotope
18. What is the mass number of an atom that consists of 20 protons, 20 neutrons, and 18 electrons?
- A) 18   B) 20   C) 38   D) **40**
19. An atom that contains 8 protons, 8 electrons, and 9 neutrons has
- A) an atomic number of 9  
B) an atomic number of 16  
C) **a mass number of 17**  
D) a mass number of 25
20. Compared to an atom of C-12, an atom of C-14 has a greater
- A) number of electrons  
B) number of protons  
C) atomic number  
D) **mass number**
21. Atoms of different isotopes of the same element differ in their total number of
- A) electrons                      B) **neutrons**  
C) protons                         D) valence electrons
22. Which particle has two neutrons?
- A)  ${}_{1}^{1}\text{p}$    B)  ${}_{1}^{1}\text{H}$    C)  ${}_{1}^{2}\text{H}$    D)  ${}_{2}^{4}\text{He}$

# Atomic Structure And Periodicity Practice Test Answers

**AW Rasmussen**



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