

CHEMISTRY

OPEN

PERIODIC TABLE OF THE ELEMENTS

Diagram showing the periodic table of elements, color-coded by groups. Hydrogen (H) is highlighted with a callout box:

- 1.008** — atomic weight
- H** — atomic symbol
- 1** — atomic number



ATOMIC STRUCTURE

Atomic Number, Z , of an element is the number of protons in the nucleus. For neutral atoms, $Z = +$ the number of electrons. The mass number, A , is the sum of the number of protons and the number of neutrons. Isotopes of an element have the same Z , different A . The atomic mass is the weighted average of isotopes. See also: Atomic Mass, Atomic Weight, Crystal Structure of Solids, Isotopes for a given element.

ATOMIC QUANTUM NUMBERS AND ORBITALS

Principal Quantum Number (n): $n = 1, 2, 3, \dots$
Angular Momentum Quantum Number (l): $l = 0, 1, 2, \dots, n-1$
Magnetic Quantum Number (m_l): $m_l = -l, -(l-1), \dots, 0, \dots, (l-1), l$
Spin Quantum Number (m_s): $m_s = +\frac{1}{2}, -\frac{1}{2}$

NUCLEAR CHEMISTRY

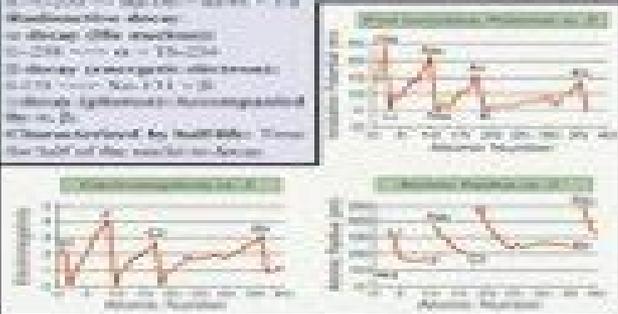
Nuclear reactions alter the number, particles present, or chemical behavior of nuclei and have energy. Energy (E) is related to mass (m) by the Einstein equation, $E = mc^2$.
 Number of nuclei that decay, N , under a constant λ is $N = N_0 e^{-\lambda t}$.
 Half-life of a nucleus, $T_{1/2}$, is $T_{1/2} = \frac{\ln 2}{\lambda}$.
 Types of processes:
 - **Radioactive decay:** α decay, β decay, γ decay, δ decay, ϵ decay, ζ decay, η decay, θ decay, ι decay, κ decay, λ decay, μ decay, ν decay, ξ decay, \omicron decay, π decay, ρ decay, σ decay, τ decay, υ decay, ϕ decay, χ decay, ψ decay, ω decay, η decay, θ decay, ι decay, κ decay, λ decay, μ decay, ν decay, ξ decay, \omicron decay, π decay, ρ decay, σ decay, τ decay, υ decay, ϕ decay, χ decay, ψ decay, ω decay.

MANY-ELECTRON ATOMS

Orbitals and quantum numbers for many-electron atoms are used to describe all atoms, molecules, ions, and solids. A many-electron atom has many discrete energy levels and orbital properties. The main point is that orbitals have different energies, the filling of levels is guided by the Aufbau Principle.

n	l	m_l	m_s	Orbitals
1	0	0	$\pm \frac{1}{2}$	1 orbital spherical
2	0	0	$\pm \frac{1}{2}$	1 orbital spherical
2	1	0, ± 1	$\pm \frac{1}{2}$	3 orbitals dumbbell
3	0	0	$\pm \frac{1}{2}$	1 orbital spherical
3	1	0, ± 1	$\pm \frac{1}{2}$	3 orbitals dumbbell
3	2	0, $\pm 1, \pm 2$	$\pm \frac{1}{2}$	5 orbitals dumbbell

Atomic Configuration: Shows orbital occupancy. Pauli Exclusion Principle: Each orbital has a unique set of quantum numbers. An orbital may hold up to two electrons, one with spin up, $m_s = +\frac{1}{2}$, and one with spin down, $m_s = -\frac{1}{2}$.
Energy Levels: Shows energy levels of an atom. The energy levels are discrete and approach a continuum at high energy. The energy levels are labeled by the principal quantum number, n , and the angular momentum quantum number, l . The energy levels are labeled by the principal quantum number, n , and the angular momentum quantum number, l . The energy levels are labeled by the principal quantum number, n , and the angular momentum quantum number, l .



Chemistry Quickstudy Reference Guides Academic

J Spring



Chemistry Quickstudy Reference Guides Academic:

This book delves into Chemistry Quickstudy Reference Guides Academic. Chemistry Quickstudy Reference Guides Academic is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Chemistry Quickstudy Reference Guides Academic, encompassing both the fundamentals and more intricate discussions.

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- Chapter 1: Introduction to Chemistry Quickstudy Reference Guides Academic
- Chapter 2: Essential Elements of Chemistry Quickstudy Reference Guides Academic
- Chapter 3: Chemistry Quickstudy Reference Guides Academic in Everyday Life
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- Chapter 5: Conclusion

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