

Equilibrium Constant - Worksheet 1_ANSWERS

1. Write the equilibrium expression for the following reactions:



$$K = \frac{[U]^u [V]^v}{[W]^w [X]^x}$$



$$K = \frac{[H_2][CO_2]}{[H_2O][CO]}$$



$$K = \frac{[Cl_2][CO]}{[COCl_2]}$$



$$K = \frac{[HCl]^2}{[H_2][Cl_2]}$$



$$K = \frac{[NO][CO_2]}{[NO_2][CO]}$$



$$K = \frac{[Zn^{2+}]}{[Ag^+]^2}$$

*Note: solids are not included



$$K = \frac{[H_2][C_2H_4]}{[C_2H_6]}$$

2. Consider the following equilibrium reaction:



The K_{eq} for this reaction at 25°C is 1.02. At equilibrium the concentration of HBr is 0.50 mol/L. Assuming H_2 and Br_2 are present in equal amounts, calculate the concentration of H_2 at equilibrium.

$$[H_2] = 0.50 \text{ mol/L}$$

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General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

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The Science Teacher, 2008. SCC Library has 1964 cur. **Illinois Chemistry Teacher**, 1992. Current Index to Journals in Education, Semi-Annual Cumulation, January-June, 1977. Educational Resources Information Center Staff, 1977-09. **Current Index to Journals in Education**, 1977. *Biology*, 1999. Chemical Equilibrium. William Guenther, 1975-12. The present work is designed to provide a practical introduction to aqueous equilibrium phenomena for both students and research workers in chemistry, biochemistry, geochemistry, and interdisciplinary environmental fields. The pedagogical strategy I have adopted makes heavy use of detailed examples of problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature. The procedure starts with mathematically complete equations that will provide valid solutions of equilibrium problems instead of the traditional approach through approximate concentrations and idealized infinite dilution.

assumptions There is repeated emphasis on the use of corrected conditional equilibrium constants and on the checking of numerical results by substitution in complete equations and or against graphs of species distributions Graphical methods of calculation and display are used extensively because of their value in clarifying equilibria and in leading one quickly to valid numerical approximations The coverage of solution equilibrium phenomena is not however exhaustively comprehensive Rather I have chosen to offer fundamental and rigorous examinations of homogeneous step equilibria and their interactions with solubility and redox equilibria Many examples are worked out in detail to demonstrate the use of equilibrium calculations and diagrams in various fields of investigation

The Software Encyclopedia, 1997 Synthesis and Characterization of Nitric Oxide-releasing Agents/polymers for Biomedical Applications Melissa May Batchelor, 2004 The Stock Exchange Handbook, 1983 **Chemical Equilibrium** Robert Blanchard Fischer, Dennis G. Peters, 1970

Equilibrium Calculations John T. Donoghue, 1971 InfoWorld, 1982-08-02 InfoWorld is targeted to Senior IT professionals Content is segmented into Channels and Topic Centers InfoWorld also celebrates people companies and projects Popular Electronics, 1982 **Chemical Equilibria in Solution** M. J. Blandamer, 1992 Popular Computing, 1982

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