

# Atomic structure and the periodic table

## Atomic structure



protons  
neutrons  
electrons

atoms have no overall charge

particle	relative mass	relative charge
proton	1	+1
neutron	1	0
electron	very small	-1

## elements

elements - substances made up of atoms with the same atomic number

isotopes of an element - atoms with the same number of protons but different numbers of neutrons

relative atomic mass - the average mass number for an element

## Equations, mixtures and chromatography

### chemical formulas and equations

chemical formula - shows the proportion of atoms of each element in a compound

chemical equation - shows the overall change in a reaction

example:

reactants

products

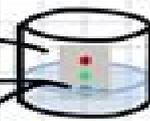
word equation: methane + oxygen → carbon dioxide + water

symbol equation:  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$

chromatogram

insoluble components stay on the base line

components separate as they move up the paper



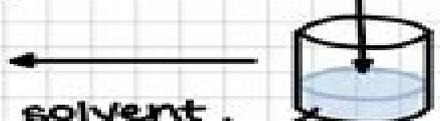
submerge it



### paper chromatography

filter paper

spot of ink



solvent

beaker

## nuclear symbols

mass number = total number of protons and neutrons in an atom



atomic number = number of protons in an atom

## compounds

compound - a substance formed from two or more elements chemically bonded together in fixed proportions

molecule - particle containing two or more non-metal atoms bonded covalently. Molecules can be elements or compounds

## mixtures

mixtures - substances made up of different elements or compounds that are not chemically bonded to each other

mixtures can be separated using physical methods

# Atomic Structure Guided Notes Answers

**Jessica J Manson**



## Atomic Structure Guided Notes Answers:

**Edexcel AS/A Level Year 1 Chemistry Student Guide: Topics 1-5** George Facer, Rod Beavon, 2015-10-09 Exam Board Edexcel Level AS A level Subject Chemistry First Teaching September 2015 First Exam June 2016 Reinforce students understanding throughout their course with clear topic summaries and sample questions and answers to help your students target higher grades Written by experienced examiners George Facer and Rod Beavon our Student Guides are divided into two key sections content guidance and sample questions and answers Content guidance will Develop students understanding of key concepts and terminology this guide covers topics 1 5 atomic structure and the periodic table bonding and structure redox 1 inorganic chemistry and the periodic table formulae equations and amounts of substance Consolidate students knowledge with knowledge check questions at the end of each topic and answers in the back of the book Sample questions and answers will Build students understanding of the different question types so they can approach questions from topics 1 5 with confidence Enable students to target top grades with sample answers and commentary explaining exactly why marks have been awarded

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Study Guide and Student Solutions Manual to Accompany Physics for Scientists and Engineers, by Serway Raymond A. Serway, Steven Van Wyk, 1996

**A Reader's Guide to Contemporary Literature** William Swan Sonnenschein, 1895

*A Reader's Guide to the Choice of the Best Available Books (about 50,000) in Every Department of Science, Art & Literature, with the Dates of the First & Last Editions, & the Price, Size & Publisher's Name of Each Book* William Swan Sonnenschein, 1901

*Study Guide, Student Solutions Manual* John R. Gordon, 1998

**The Brass World and Platers Guide**, 1922

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**The Official Guide for GMAT Review**, 1988

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**The Chemical News and Journal of Industrial Science**, 1906

**Study Guide/Selected Solutions Manual** Julie R. Frentrup, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

The Praxis Series Official Study Guide: Middle School: Science,

**Isis** George Sarton, 1924 Brief table of contents of vols I XX in v 21 p 502 618

**Instructor's Guide for Introductory Chemistry in the Laboratory** James F. Hall, 1996

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**Reader's Guide to Periodical Literature Supplement**, 1922

Research in Education, 1973

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,1920

## **Atomic Structure Guided Notes Answers** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Atomic Structure Guided Notes Answers**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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