

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the structure corresponding to the term in the vocabulary list.

- | | |
|-------------------------|---------------------------|
| _____ mole | _____ percent composition |
| _____ empirical formula | _____ empirical formula |
| _____ molar mass | _____ molar mass |
| _____ molar mass | _____ molar mass |
10. _____ is a description of the relative amounts of each element in a compound. _____ percent composition
11. _____ is the lowest whole number ratio of the atoms of the elements in a compound. _____ empirical formula
12. _____ is the mass of one mole of any gas at STP. _____ molar mass
13. _____ is the mass of one mole of a compound. _____ molar mass
14. _____ is the amount representing 6.02×10^{23} representative particles of a substance. _____ mole
15. _____ is 6.02×10^{23} particles. _____ molar mass
16. _____ is the mass of one mole of a substance at STP. _____ molar mass
17. _____ is the mass of one mole of a gas at STP. _____ molar mass
18. _____ is the mass of one mole of a gas at STP. _____ molar mass

Part 2: Section Concept Review

10.1 The Mole: A Fundamental

Complete the following activities to demonstrate knowledge of the terms and your understanding of the concepts involved with a mole of a substance. Complete each activity on your own.

1. _____ is the mass of one mole of a substance. _____ is the mass of one mole of a substance. _____ is the mass of one mole of a substance. _____ is the mass of one mole of a substance.

2. _____ is the mass of one mole of a compound. _____ is the mass of one mole of a compound. _____ is the mass of one mole of a compound. _____ is the mass of one mole of a compound.

Practice Problems

1. _____ is the mass of one mole of a substance. _____ is the mass of one mole of a substance.

2. _____ is the mass of one mole of a substance. _____ is the mass of one mole of a substance.

3. _____ is the mass of one mole of a substance. _____ is the mass of one mole of a substance.

4. _____ is the mass of one mole of a substance. _____ is the mass of one mole of a substance.

5. _____ is the mass of one mole of a substance. _____ is the mass of one mole of a substance.

6. _____ is the mass of one mole of a substance. _____ is the mass of one mole of a substance.

Challenge

Chapter 10 Chemical Quantities Guided Reading Answer Key

M Mosston



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